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a net ionic equation is the most accurate representation of the actual chemical process that occurs writing these equations requires a familiarity with solubility rules acid base reactivity

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apr 25 2017 this is understood when you consider the two chemicals appear unchanged before and after the reaction therefore the net ionic equation is $2\text{Cl}^- \text{aq} + 2\text{Ag}^+ \text{aq}$

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ions that remain essentially unchanged during a reaction are called spectator ions this means you can ignore them when you write the ionic equation you only need to model how the solid

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1 $\text{K}^+ \text{aq} + \text{Al}^{3+} \text{aq} + \text{PO}_4^{3-} \text{aq} \rightarrow \text{AlPO}_4 \text{s}$ 3 $\text{K}^+ \text{aq} + 3\text{PO}_4^{3-} \text{aq} \rightarrow 3\text{K}_3\text{PO}_4 \text{s}$ net ionic $\text{Al}^{3+} \text{aq} + \text{PO}_4^{3-} \text{aq} \rightarrow \text{AlPO}_4 \text{s}$ 2 $\text{Fe}^{2+} \text{aq} + 3\text{Mg} \text{s}$

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