

### 18.4 Write the following in words:

#### 18.4.1 Write in words:

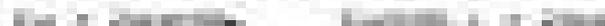
1. The concentration of quantities in chemical equations is called \_\_\_\_\_  
 (a) percent concentration (b) percent yield  
 (c) molar fraction (d) molar concentration

2. Write in words the following chemical equations in concentration terms (molar concentration):



- (a) 2 g of  $\text{H}_2$  and 16 g of  $\text{O}_2$  (b) 2 g of  $\text{H}_2$  (c) 16 g of  $\text{O}_2$  (d) 2 g of  $\text{H}_2$  and 16 g of  $\text{O}_2$   
 (e) 2 moles of  $\text{H}_2$  and 1 mole of  $\text{O}_2$  (f) 2 moles of  $\text{H}_2$  (g) 16 moles of  $\text{O}_2$  (h) 2 moles of  $\text{H}_2\text{O}$

3. In a gaseous mixture hydrogen sulphide (molar fraction 0.2) and sulfur dioxide, 0.5 g of  $\text{H}_2\text{S}$  are separated by law fractional distillation, and only 0.1 g of  $\text{H}_2\text{S}$  are obtained. What is the percent yield of  $\text{H}_2\text{S}$  in this reaction?



- (a) 20% (b) 25% (c) 20.5% (d) 20.1% (e) 20.2%

4. In a gas mixture 200 g of  $\text{CO}_2$  and 200 g of  $\text{CO}$  are separated by law fractional distillation. What is the percent yield of  $\text{CO}_2$  in this reaction?

- (a) 10% (b) 15% (c) 20% (d) 25%

5. Write in words the following chemical equations in concentration terms:



- (a) 1 g (b) 1 g (c) 1 g (d) 1 g (e) 1 g

6. A mixture of 100 g of  $\text{H}_2$  and 100 g of  $\text{O}_2$  is separated by law fractional distillation. What is the percent yield of  $\text{H}_2$  in this reaction?



- (a) 100 g (b) 100 g (c) 100 g (d) 100 g (e) 100 g

7. In a gas mixture 100 g of  $\text{CO}_2$  and 100 g of  $\text{CO}$  are separated by law fractional distillation. What is the percent yield of  $\text{CO}_2$  in this reaction?



# Multiple Choice Stoichiometry Test

**Y Pai**

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