

Voltaic Cells

Label the parts for the following cell:



By connecting to the voltmeter, describe a spontaneous redox reaction. (10 marks)

- a. anode: _____
- b. cathode: _____
- c. anode: _____
- d. cathode: _____
- e. half-reaction (oxidation) at anode: _____
- f. half-reaction (reduction) at cathode: _____
- g. net redox (balanced) equation: _____
- h. salt: _____
- i. better metals will always reduce/better metals will always oxidize, except for better arguments with half-reactions for specific metals

Voltaic Cells Study Guide

Yijin Wang



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