

18.4 Acid-Base Chemistry: Theory

Multiple-Choice Questions

1. Which combination of quantities are dimensional quantities? (correct)
- (A) pressure, concentration, pH, amount of substance
 (B) volume, molarity, pH, amount of substance
 (C) volume, molarity, pH, concentration of ions
 (D) amount of substance, pH, concentration of ions, concentration of substance

(A) $n = 1000$, $m = 1000$

(B) $10 \text{ g H}_2\text{O}$, $1 \text{ mol H}_2\text{O}$, $1 \text{ g H}_2\text{O}$, $1 \text{ mol H}_2\text{O}$, $10 \text{ g H}_2\text{O}$, $1 \text{ mol H}_2\text{O}$
 (C) $1 \text{ mol H}_2\text{O}$, $1 \text{ mol H}_2\text{O}$

2. An acid-base indicator has a pK_a of 4.5. In a solution of pH 5.5, the indicator is present in the form of its conjugate base. What is the ratio of the concentration of the conjugate base to the concentration of the indicator? (correct)

(A) 10^{10}
 (B) 10^9
 (C) 10^8
 (D) 10^7

(E) 10^6
 (F) 10^5
 (G) 10^4
 (H) 10^3
 (I) 10^2
 (J) 10^1

3. An acid-base indicator has a pK_a of 4.5. In a solution of pH 5.5, the indicator is present in the form of its conjugate base. What is the ratio of the concentration of the conjugate base to the concentration of the indicator? (correct)

(A) 10^1
 (B) 10^2
 (C) 10^3
 (D) 10^4

4. Which of the following is a weak acid? (correct)



(E) H_2O
 (F) H_2S
 (G) H_2SO_3
 (H) H_2SO_4

5. A solution of an acid has a pH of 2.5. The concentration of the acid is 0.1 mol/L. What is the pK_a of the acid? (correct)

(A) 1.5
 (B) 2.5
 (C) 3.5
 (D) 4.5

(E) 5.5
 (F) 6.5
 (G) 7.5
 (H) 8.5

6. An acid-base indicator has a pK_a of 4.5. In a solution of pH 5.5, the indicator is present in the form of its conjugate base. What is the ratio of the concentration of the conjugate base to the concentration of the indicator? (correct)

(A) 10^1
 (B) 10^2
 (C) 10^3
 (D) 10^4

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