

# Mechanic Diesel – Semester 2 Module 1 - Diesel Engine Overview

## ITI Mechanic Diesel Trade Theory

## ITI मैकेनिक डीजल

01

What is the type of engine? | इंजन का प्रकार क्या है?



- A : In-line engine | इन-लाइन इंजन
- B : V engine | वी इंजन
- C : Opposed engine | अपोजिट इंजन
- D : Radial engine | रेडियल इंजन

What is the name of the stroke? | स्ट्रोक का नाम क्या है?



- A : Compression stroke | संपीड़न स्ट्रोक
- B : Power stroke | पावर स्ट्रोक
- C : Suction stroke | सक्शन स्ट्रोक
- D : Exhaust stroke | निष्कास स्ट्रोक

What is the type of engine? | इंजन का प्रकार क्या है?



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**Clemens Wendtner**



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 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) \rightarrow \text{C}_2\text{H}_2(\text{g}) + \frac{5}{2} \text{O}_2(\text{g})$   $\Delta h = 226.7 \text{ kJ}$   
 $2 \text{ H}_2(\text{g}) + \text{C}_2\text{H}_6(\text{g}) \rightarrow \Delta h = 94.5 \text{ kJ}$   
 $\text{H}_2\text{O}(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \Delta h = 71.2 \text{ kJ}$   
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{g})$   $\Delta h = 283 \text{ kJ}$

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 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 2 \text{H}_2\text{O}(\text{l})$   $\Delta h = 1411 \text{ kJ}$  mole  
 $\text{C}_2\text{H}_6(\text{g}) + 7/2 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{l})$   $\Delta h = 1560 \text{ kJ}$  mole  
 $\text{H}_2(\text{g}) + 1/2 \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$   $\Delta h = 285 \text{ kJ}$  mole

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different steps or stages in the reaction there are cohen 2016 the equations for the reactions in the experiment done are as follows 1 naoh s na aq oh aq

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web thus for reaction 9 2 in which 1 mole of  $h_3o$  and 1 mole of  $oh$  combine to form 2 moles of  $h_2o$   $\Delta h = 55.8 \text{ kJ}$  as shown note that the enthalpy change  $\Delta h_{products} - \Delta h_{reactants}$  is positive if heat is absorbed that is if  $\Delta h_{products}$

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