

14.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of species associated with
2. chemical equilibria and concentrations of ions, acids, and hydroxides and associated equilibrium constants
3. describe the properties of pH indicators

Vocabulary

- | | | |
|-------------------|------------------------------------|-------------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid indicator |
| 4. hydrolysis | 5. hydroxide ion (OH^-) | 6. equilibrium constant |
| | 7. acid indicator | 8. pH |

Key Equations

- | | |
|----------------------------------------------------------------------------|-----------------------------------------------------------------------------------------------------------|
| 1. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ (14.1) | 2. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (14.2) |
| 3. $\text{pH} = -\log [\text{H}^+]$ (14.3) | 4. $\text{pOH} = -\log [\text{OH}^-]$ (14.4) |

Part A: Completion

Use the appropriate number to identify the relationship of the species indicated that are associated in the reaction. Also indicate which species acts as the conjugate acid.

1. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (14.2) (14.1) (14.3) (14.4)
2. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (14.2) (14.1) (14.3) (14.4)
3. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (14.2) (14.1) (14.3) (14.4)
4. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (14.2) (14.1) (14.3) (14.4)
5. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (14.2) (14.1) (14.3) (14.4)
6. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (14.2) (14.1) (14.3) (14.4)
7. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (14.2) (14.1) (14.3) (14.4)
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1. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (14.2)
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