

# Voltaic Cells

Label the parts for the following cell:



By connecting to the voltmeter, describe a spontaneous redox reaction. (10 marks)

- a. metals \_\_\_\_\_
- b. metal ions \_\_\_\_\_
- c. anodes \_\_\_\_\_
- d. cathodes \_\_\_\_\_
- e. electron movement flow \_\_\_\_\_
- f. displacement overall movement \_\_\_\_\_

- g. half-cell reduction equation \_\_\_\_\_
- h. Zn \_\_\_\_\_

i. Write overall cell equation with phases.  $\text{Zn(s)} + \text{Cu}^{2+}(\text{aq}) \rightarrow \text{Zn}^{2+}(\text{aq}) + \text{Cu(s)}$  (10 marks)

\_\_\_\_\_

# Voltaic Cells Study Guide

**Ian Pickup**



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