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with 1200 ml of 2m hno3 if the final density of solution is same as pure water then molarity of the iron in resulting solution

which decides the nature of the above solution is 1 0 25m 2 0 5m 3 0 888m 4 none of

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concentration units and they are used interchangeably

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0 10 0 0585 correct answer b 1 explanation molality
$$l = \frac{1000 \text{ m}}{1000d}$$

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pressure and temperature changes concentrations expressed in molality are used

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the molar concentration of the solute is sometimes abbreviated by putting square brackets around the chemical formula of

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web apr 7 2014 molarity moles of solute liters of solution 8.4 2.2 a first convert 250 ml to liters 250/1000 = 0.25 then calculate molarity 5 moles / 0.25 liters = 20 M 3. c a solution with molarity 2 requires 2 M of NaOH per liter so 4 x 2.8 M = 11.2 M a solution of molarity 1.5 M requires 1.5 mol of Na to every litre of solvent

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web answer the correct option is c explanation the molarity of a mixture mix can be calculated using the following formula $M_{mix} = \frac{m_1}{V_1} + \frac{m_2}{V_2} + \frac{m_3}{V_3} + \dots$ 1.5 x 480 / 1.2 x 520 = 480 / 520 = 0.923 1000 / 1.344 = 744.045 with increase in temperature which of these changes a molality b fraction of solute present in water

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