

18.4 Acid-Base Chemistry: Theory

Multiple-Choice:

1. Which combination of quantities are dimensional quantities? (correct)
- (A) pressure, concentration, pH, amount of substance
 (B) volume, molarity, pH, amount of substance
 (C) volume, molarity, pH, concentration of ions
 (D) pH, amount of substance, concentration of ions, concentration of substance

2. Which of the following is FALSE? (correct)
- (A) 10^{-14} mol/L of H^+ is 10^{-14} mol/L of OH^- .
 (B) 10^{-14} mol/L of H^+ is 10^{-14} mol/L of OH^- .
 (C) 10^{-14} mol/L of H^+ is 10^{-14} mol/L of OH^- .
 (D) 10^{-14} mol/L of H^+ is 10^{-14} mol/L of OH^- .

3. In a 0.1 M acetic acid solution, the concentration of acetate ions is approximately 0.001 M. What is the pH of this solution? (correct)
- (A) 4.0
 (B) 3.0
 (C) 2.0
 (D) 1.0

4. In a 0.1 M acetic acid solution, the concentration of acetate ions is approximately 0.001 M. What is the pH of this solution? (correct)
- (A) 4.0
 (B) 3.0
 (C) 2.0
 (D) 1.0

5. Which of the following is FALSE? (correct)
- (A) 10^{-14} mol/L of H^+ is 10^{-14} mol/L of OH^- .
 (B) 10^{-14} mol/L of H^+ is 10^{-14} mol/L of OH^- .
 (C) 10^{-14} mol/L of H^+ is 10^{-14} mol/L of OH^- .
 (D) 10^{-14} mol/L of H^+ is 10^{-14} mol/L of OH^- .

6. A solution of acetic acid and its conjugate base is prepared with a pH of 4.0. What is the ratio of the concentration of the conjugate base to the concentration of the acid? (correct)
- (A) 10:1
 (B) 1:10
 (C) 1:1
 (D) 10:1

7. In a 0.1 M acetic acid solution, the concentration of acetate ions is approximately 0.001 M. What is the pH of this solution? (correct)
- (A) 4.0
 (B) 3.0
 (C) 2.0
 (D) 1.0

8. In a 0.1 M acetic acid solution, the concentration of acetate ions is approximately 0.001 M. What is the pH of this solution? (correct)
- (A) 4.0
 (B) 3.0
 (C) 2.0
 (D) 1.0

Multiple Choice Stoichiometry Test

Chao Zhang



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