

# CAMPBELL'S BIOLOGY 10TH EDITION CH 4, 5, 6 EXAM STUDY GUIDE QUESTIONS WITH CORRECT ANSWERS

Tissues - Answer-A group of specialized cells that perform a particular function

Plant cells contain: - Answer-Cell wall; chloroplasts; central vacuole

Stanley Miller's experiment in 1953 confirmed: - Answer-Organic molecules can be made from non living molecules

Abiotic synthesis - Answer-Creating organic molecules from non living materials

Carbon - Answer-Can make four covalent (unusual) bonds to fill outer shell to 8

Hydrogen valence - Answer-=1

Oxygen valence - Answer-=2

Nitrogen valence - Answer-=3

Carbon valence - Answer-=4

C<sub>2</sub>H<sub>6</sub> - Answer-H H

||

H-C-C-H

||

H H

Double bonds - Answer-Share the carbons

Triple bonds - Answer-are possible between two carbons

Structural Isomer - Answer-Same molecular formula, different structure-makes them different chemicals

Cis-trans Isomer - Answer-Carbon to carbon double bonds

Cis-Isomer - Answer-Carbon bonds to other molecules from the same side; nature prefers

Trans-Isomer - Answer-Carbon bonds to other molecules one on either side

# Campbell Guide Answers

**J Dewey**



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web in today s experiment you will determine the solubility product ksp of calcium hydroxide ca oh 2 by measuring the concentration of ca oh 2 in a saturated solution calcium hydroxide is a sparingly soluble salt that dissolves according to the following reaction the solubility product expression for this reaction is

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web expert verified chem 182 determining the ksp of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water a saturated aqueous solution of ca oh is represented in equation form as shown below ca oh s ca aq 2oh aq the solubility product expression describes in mathematical terms the equilibrium

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web use the concentration of calcium hydroxide to calculate the concentration of calcium ions and hydroxide ions in the solution ca oh 2 à ca 2 2oh 5 x 10 2 m calcium ions 5 x 10 2 m 2 0 m hydroxide ions calculate calcium hydroxide ksp using the concentrations of calcium ions and hydroxide ions

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web madison bowe elana keller april 15 2014 introduction the objective of this lab was to measure the ksp of a slightly soluble basic salt and how solubility changes with temperature ksp is another form of equilibrium expression that deals with solubility the reaction that will take place is ca oh 2 ca2 2 oh

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oh 2 m ca in saturated ca oh 2 m ksp of ca oh 2 average ksp of ca oh 2

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web the chemical equation used was  $\text{Ca(OH)}_2 \text{ (s)} \rightleftharpoons \text{Ca}^{2+} \text{ (aq)} + 2\text{OH}^{-} \text{ (aq)}$  the solubility product was determined to be  $K_{sp} \text{ Ca(OH)}_2$  from the titration which was made by dripping a concentrated solution of calcium hydroxide into hydrochloric acid the calcium hydroxide was filtered before starting the titration

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web jul 12 2023 because the concentration of a pure solid such as  $\text{Ca}_3\text{PO}_4\text{ (s)}$  is a constant it does not appear explicitly in the equilibrium constant expression the equilibrium constant expression for the dissolution of calcium phosphate is therefore  $K_{sp} \text{ Ca}_3\text{PO}_4\text{ (s)} = [\text{Ca}^{2+}]^3 [\text{PO}_4^{3-}]^2$

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web therefore it was determined through error calculations and ease of use that the most reliable valid method to use to get the closest ksp value to the theoretical ksp for ca was the determination through ph method 2 introduction calcium hydroxide is a strong base soluble in water

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