

KEY

1 26. Which is the stronger base, HCO_3^- or HCOO^- ? (1 mark) HCO_3^-

27. Classify each of the following as: a strong acid (SA), weak acid (WA), strong base (SB), weak base (WB) or a spectator ion (S). (10 marks)

10

- | | | | |
|--------------------------------|-----------|-----------------------------|-----------|
| a) F^- | <u>WB</u> | f) Cl^- | <u>S</u> |
| b) HIO_3 | <u>WA</u> | g) NH_3 | <u>WB</u> |
| c) NO_3^- | <u>S</u> | h) O^{2-} | <u>SB</u> |
| d) HClO_4 | <u>SA</u> | i) CH_3COOH | <u>WA</u> |
| e) $\text{C}_2\text{O}_4^{2-}$ | <u>WB</u> | j) ClO_4^- | <u>S</u> |

28. What is the $[\text{OH}^-]$ in a solution made by adding 0.060 moles of calcium oxide to 500.0 mL water? Be careful! (2 marks)

2

$$M = \frac{\text{mol}}{L} = \frac{0.060 \text{ mol CaO}}{0.500 \text{ L}} = 0.12 \text{ M} = [\text{CaO}]$$

$$0.12 \text{ CaO} \rightarrow \text{Ca}^{2+} + 0.12 \text{ O}^{2-} \Rightarrow \text{O}^{2-} + \text{H}_2\text{O} \rightarrow 2\text{OH}^-$$

Answer $[\text{OH}^-] = 0.24 \text{ M}$

29. What is the $[\text{H}_3\text{O}^+]$ in a solution made by adding 0.020 moles of nitric acid to 500.0 mL of water? (2 marks)

2

$$[\text{HNO}_3] = \frac{0.020 \text{ mol}}{0.500 \text{ L}} = 0.040 \text{ M}$$

SA so $[\text{H}_3\text{O}^+] = [\text{acid}] = 0.040 \text{ M}$

Answer $[\text{H}_3\text{O}^+] = 0.040 \text{ M}$

30. If 0.10 M HSO_3^- is mixed with 0.10 M HC_2O_4^- , which species will donate a proton? (1 mark)

1

Answer HC_2O_4^-

31. If 0.10 M HSO_4^- is mixed with 0.10 M $\text{HC}_2\text{H}_3\text{O}_2^-$, which species will donate a proton? (1 mark)

1

Answer HSO_4^-

$\frac{17}{17}$

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ML Morrison



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