

18.4 Molecular Structure: Mass

Multiple Choice

1. The molar masses of quantities in chemical equations are based _____
 (A) percent composition (B) percent yield
 (C) molar mass (D) molar volume
2. Percent of oxygen in $\text{CaSO}_4 \cdot \frac{1}{2}\text{H}_2\text{O}$ is approximately (to two significant figures) _____

(A) 18% (B) 23% (C) 28% (D) 33%

3. In a gas the other reactions represent copper metal and silver chloride, 2.5 g of Ag are completely reacted by the reaction, and only 0.5 g of Ag metal is left. What is the percent yield of silver in this reaction?

(A) 20% (B) 40% (C) 60% (D) 80%

4. In this reaction $2\text{CO(g)} + \text{O}_2\text{(g)} \rightarrow 2\text{CO}_2\text{(g)}$, which is the correct number of moles of oxygen that reacts with 2 moles of CO ?

(A) 1 (B) 2 (C) 3 (D) 4

5. Which is the correct number of moles in the following equation?



(A) 2 (B) 1 (C) 3 (D) 4

6. Aluminum reacts with HCl to produce hydrogen gas. If 1.5 g of aluminum reacts with 100 mL of 1.0 M HCl, what mass of H_2 gas is produced?



(A) 0.05 g (B) 0.15 g (C) 0.45 g (D) 1.35 g (E) 2.25 g

7. In a full combustion, structure-molecular weight $\text{C}_2\text{H}_5\text{CO}_2\text{C}_2\text{H}_5$ what carbon atom is oxidized, C_1 or C_2 ? After this reaction was completed, resulting water percent was 50% relative to the reaction mixture. What was the limiting reagent?



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G Thomas

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