

14.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations and pH values and vice versa for a given pH.
3. Calculate the pH of weak acid solutions.

Vocabulary

- | | | |
|-------------------|--------------------------------|--------|
| 1. autoionization | 2. conjugate acid-base pair | 3. pH |
| 4. hydronium ion | 5. hydronium ion concentration | 6. pOH |
| | 7. hydroxide ion | |

Key Equations

- | | |
|---|---|
| 1. $\text{pH} = -\log[\text{H}_3\text{O}^+]$ (Eq. 14.1) | 2. $\text{pOH} = -\log[\text{OH}^-]$ (Eq. 14.2) |
| 3. $\text{pH} + \text{pOH} = \text{pK}_w$ (Eq. 14.3) | 4. $\text{pH} + \text{pOH} = 14.00$ (Eq. 14.4) |

Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair indicated in the second step of the acid-base reaction given in the following examples.

1. What product is the _____ of the hydroxide ion and the hydronium ion? _____
2. The conjugate base of the hydronium ion is the _____ of the hydronium ion. _____
3. _____ is the conjugate acid of the hydroxide ion. _____
4. _____ is the conjugate base of the hydronium ion. _____
5. _____ is the conjugate acid of the hydroxide ion. _____
6. _____ is the conjugate base of the hydronium ion. _____
7. _____ is the conjugate acid of the hydroxide ion. _____
8. _____ is the conjugate base of the hydronium ion. _____
9. _____ is the conjugate acid of the hydroxide ion. _____
10. _____ is the conjugate base of the hydronium ion. _____

1. _____
2. _____
3. _____
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10. _____

19 2 Hydrogen Ions And Acidity Answers

James Small



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