



Name

Date

## Student Exploration: Covalent Bonds

**Directions:** Follow the instructions to go through the simulation. Respond to the questions and prompts in the orange boxes.

**Vocabulary:** covalent bond, diatomic molecule, Lewis diagram, molecule, noble gases, nonmetal, octet rule, shell, valence, valence electron

**Prior Knowledge Questions** (Do these BEFORE using the Gizmo.)

1. There are eight markers in a full set, but Fiona and Frank each only have seven markers. Fiona is missing the red marker, and Frank is missing the blue marker.

What can they do so that each has a full set of markers?

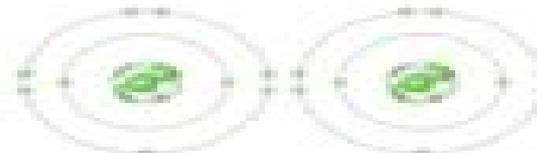
Frank should give the red marker to Fiona, and Fiona should give the blue marker to Frank.

2. Otto and Olivia each have six markers. Otto is missing the purple and green markers, and Olivia is missing the black and brown markers. What can they do so that each has a full set?

Otto should give the black and brown markers to Olivia, and Olivia should give Otto the purple and green markers.

### Gizmo Warm-up

Just like the students described above, **nonmetal** atoms can share electrons. As you will see in the Covalent Bonds Gizmo, atoms form bonds in this way.



To begin, check that **Fluorine** is selected from the **Select a substance** menu. Click **Play** (  ) to see the electrons orbiting the nucleus of each atom.

1. The outermost electrons in each atom are called **valence electrons**. How many valence electrons does each fluorine atom have?

2. Click **Pause** (  ). Drag a valence electron from the left atom to the right atom. Click **Play**.

What happens?

3. Click **Pause**, drag a valence electron from the right atom to the left, and then click **Play**.

# Section 2 nuclear Chemistry Study Guide Key

**N Colangelo**



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**Study Guide with Answers to Selected Problems** H. Stephen Stoker, Danny V. White, 2003-04 This supplement includes for each chapter a brief overview activities and practice problems to reinforce skills and a practice test The answers section includes answers for all odd numbered end of chapter exercises *Chemistry, Student Study Guide* John A. Olmsted, Gregory M. Williams, 2005-02-02 100% Pure Chemical Understanding Every morning many of us are energized by a cup of coffee Imagine if you were as energized by understanding the chemistry in your morning cup from the coffee trees which fill red coffee berries with caffeine and a variety of other chemical substances to the feathery crystals formed by the caffeine molecules to the decaffeinating machines which use liquid solvents to remove this stimulant from some of the beans Now that s real chemical understanding Olmsted and Williams Fourth Edition of Chemistry focuses on helping you see and think about the world and even your coffee as a chemist This text helps you understand how chemical phenomena are governed by what happens at the molecular level apply critical thinking skills to chemical concepts and problems and master the basic mathematical techniques needed for quantitative reasoning You ll see the world as chemists do and learn to appreciate the chemical processes all around us A Fourth Edition with a lot of new perks Revisions include a new early energy chapter revised coverage of bonding expanded coverage of intermolecular forces and increased coverage of multiple equilibria including polyprotic acids New pedagogy strengthens students critical thinking and problem solving skills Visual Summaries at the end of each chapter use molecular and diagrammatic visual elements to summarize essential skills concepts equations and terms eGrade Plus provides an integrated suite of teaching and learning resources including a complete online version of the text links between problems and relevant sections in the online text practice quizzes the Visual Tutor Interactive LearningWare problems and lab demos as well as homework management and presentation features for instructors

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