

The curve in Figure 16.11 of the text shows how the vibrational heat capacity depends on temperature. Note that even when the temperature is only slightly above the vibrational temperature, the heat capacity is close to its equipartition value.

The total heat capacity of a molecular substance is the sum of each contribution (Figure 16.12 of the text). When equipartition is valid (when the temperature is well above the characteristic temperature of the mode $T \gg \theta_M$) we can estimate the heat capacity by counting the numbers of modes that are active. In gases, all three translational modes are always active and contribute $\frac{3}{2}R$ to the molar heat capacity. If we denote the number of active rotational modes by ν_R^\ddagger (so for most molecules at normal temperatures $\nu_R^\ddagger = 2$ for linear molecules, and 3 for non-linear molecules), then

$$C_{V,m} = \frac{1}{2}R(3 + \nu_R^\ddagger + \nu_V^\ddagger)$$

The equipartition value $C_{V,m}^\ddagger$ is the value of $C_{V,m}$ when all the modes are active. In the limit $T \gg \theta_M$ the equipartition value is $C_{V,m}^\ddagger = \frac{1}{2}R(3 + \nu_R^\ddagger + \nu_V^\ddagger)$. The equipartition value is the value of $C_{V,m}$ when all the modes are active. In the limit $T \gg \theta_M$ the equipartition value is $C_{V,m}^\ddagger = \frac{1}{2}R(3 + \nu_R^\ddagger + \nu_V^\ddagger)$. The equipartition value is the value of $C_{V,m}$ when all the modes are active. In the limit $T \gg \theta_M$ the equipartition value is $C_{V,m}^\ddagger = \frac{1}{2}R(3 + \nu_R^\ddagger + \nu_V^\ddagger)$.

Solutions to exercises

The characteristic temperature is given by $\theta_M = \frac{h^2}{8mk_B}$. The characteristic temperature is $\theta_M = \frac{h^2}{8mk_B}$.

$$\theta_M = \frac{h^2}{8mk_B} = \frac{(6.626 \times 10^{-34} \text{ J s})^2}{8 \times (2.016 \times 10^{-26} \text{ kg}) \times (1.381 \times 10^{-23} \text{ J K}^{-1})} = 1.92 \text{ K}$$

At 298 K, $T \gg \theta_M$, so the equipartition value is $C_{V,m}^\ddagger = \frac{1}{2}R(3 + \nu_R^\ddagger + \nu_V^\ddagger)$.

$$C_{V,m}^\ddagger = \frac{1}{2}R(3 + 2 + 0) = \frac{5}{2}R = 20.8 \text{ J K}^{-1} \text{ mol}^{-1}$$

At 100 K, $T \approx \theta_M$, so

$$C_{V,m} = \frac{1}{2}R \left(3 + \frac{2}{1 + e^{-\theta_M/T}} + \frac{0}{1 + e^{-\theta_M/T}} \right) = 18.5 \text{ J K}^{-1} \text{ mol}^{-1}$$

At 10 K, $T \ll \theta_M$, so

$$C_{V,m} = \frac{1}{2}R(3 + 0 + 0) = \frac{3}{2}R = 12.5 \text{ J K}^{-1} \text{ mol}^{-1}$$

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Table of Contents Physical Chemistry 9th Edition Solutions

1. Understanding the eBook Physical Chemistry 9th Edition Solutions
 - The Rise of Digital Reading Physical Chemistry 9th Edition Solutions
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Chemistry 9th Edition Solutions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Chemistry 9th Edition Solutions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Chemistry 9th Edition Solutions
 - Personalized Recommendations
 - Physical Chemistry 9th Edition Solutions User Reviews and Ratings
 - Physical Chemistry 9th Edition Solutions and Bestseller Lists
5. Accessing Physical Chemistry 9th Edition Solutions Free and Paid eBooks
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 - Physical Chemistry 9th Edition Solutions eBook Subscription Services

- Physical Chemistry 9th Edition Solutions Budget-Friendly Options
- 6. Navigating Physical Chemistry 9th Edition Solutions eBook Formats
 - ePub, PDF, MOBI, and More
 - Physical Chemistry 9th Edition Solutions Compatibility with Devices
 - Physical Chemistry 9th Edition Solutions Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Chemistry 9th Edition Solutions
 - Highlighting and Note-Taking Physical Chemistry 9th Edition Solutions
 - Interactive Elements Physical Chemistry 9th Edition Solutions
- 8. Staying Engaged with Physical Chemistry 9th Edition Solutions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Chemistry 9th Edition Solutions
- 9. Balancing eBooks and Physical Books Physical Chemistry 9th Edition Solutions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Chemistry 9th Edition Solutions
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Chemistry 9th Edition Solutions
 - Setting Reading Goals Physical Chemistry 9th Edition Solutions
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Chemistry 9th Edition Solutions
 - Fact-Checking eBook Content of Physical Chemistry 9th Edition Solutions
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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web yes it is possible to do pairwise comparisons in excel here are the steps to do it first you need to create a table with the items you want to compare for example if you want to compare different types of fruits you can create a table with the

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web may 31 2015 not sure you mean by without comparing each pair of cells as you have to do some work to get the result you want if the columns to compare are a and b i would put the following in cell c1 and then copy it down the entire range if a1 b1 1 then the rows that are different will have a 1 in the relevant c column

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web next pairwise compare each candidate in a row to a different candidate in a column and pairwise rank them according to who you prefer keep track using the following simple scoring system if you prefer the row candidate over the column candidate e g neha vs peter enter 1 in that cell and enter 0 in the inverse cell peter vs neha as in table 2