

### 18.4 Acid-Base Chemistry: Theory

#### Multiple-Choice Questions

1. Which combination of quantities are dimensional quantities? (correct)
- (A) pressure, concentration, pH, amount of substance  
 (B) volume, molarity, pH, amount of substance  
 (C) volume, molarity, pH, concentration of ions  
 (D) amount of substance, pH, concentration of ions, concentration of substance

(A)  $n = 1000$ ,  $m = 1000$

(B)  $10 \text{ g H}_2\text{O}$ ,  $1 \text{ mol H}_2\text{O}$ ,  $1 \text{ g H}_2\text{O}$ ,  $1 \text{ mol H}_2\text{O}$ ,  $10 \text{ g H}_2\text{O}$ ,  $1 \text{ mol H}_2\text{O}$   
 (C)  $1 \text{ mol H}_2\text{O}$ ,  $1 \text{ mol H}_2\text{O}$

2. An acid-base reaction represents a proton transfer from a stronger acid to a weaker acid. Which of the following is not correct, and why? (correct)
- (A)  $\text{HCl} + \text{H}_2\text{O} \rightarrow \text{H}_3\text{O}^+ + \text{Cl}^-$  (strong acid reacts with water to form a weaker acid)



3. An acid-base reaction involves the transfer of a proton from a stronger acid to a weaker acid. Which of the following is not correct, and why? (correct)
- (A)  $\text{HCl} + \text{H}_2\text{O} \rightarrow \text{H}_3\text{O}^+ + \text{Cl}^-$  (strong acid reacts with water to form a weaker acid)



4. Which of the following is not a property of a strong acid? (correct)



5. A solution of a weak acid and its conjugate base is prepared. The pH of the solution is 4.0. The acid dissociation constant,  $K_a$ , of the acid is  $1.0 \times 10^{-4}$ . What is the ratio of the acid to the conjugate base? (correct)



6. An acid-base reaction involves the transfer of a proton from a stronger acid to a weaker acid. Which of the following is not correct, and why? (correct)
- (A)  $\text{HCl} + \text{H}_2\text{O} \rightarrow \text{H}_3\text{O}^+ + \text{Cl}^-$  (strong acid reacts with water to form a weaker acid)



# Multiple Choice Stoichiometry Test

**G Thomas**



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