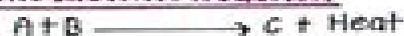


HEAT IN REACTIONS :-

(A) Exothermic Reaction



(1) Respiration



(2) Burning of natural gas



(3) Decomposition of vegetables into compost.

(B) Endothermic Reaction



- All Examples of Decomposition Reaction



reaction

REDOX REACTION The Reaction in which oxidation and Reduction occur.

Oxidation

- if a chemical (A)
 - loss of electron
 - Gains of Oxygen
 - loss of Hydrogen

It is called oxidation of A and 'A' is said to be oxidised.



Reduction

- if chemical (B)
 - Gain of e^-
 - loss of O_2 (oxygen)
 - Gains of Hydrogen

it is called Reduction of B and 'B' is said to be Reduced.



CORROSION :- When a metal is attacked by substances around it such as moisture (water vapour + oxygen), acid etc. it is said to be corrode and this process is called corrosion.

Examples of corrosion :-

(1) Rusting of iron - Hydrated Reddish Brown iron oxide (rust)

(2) Tarnishing of copper - Green

(3) Tarnishing of silver - Black

RANCIDITY :- - The taste or smell of food material containing fat/oil changes when it is left exposed to air for a long time. ✓ oxidation of fat/oil present in food material causes Rancidity.

Prevention :-

- (1) Antioxidants
- (2) Air tight Container
- (3) Bags of chips (Flushed with N_2 gas)

PHYSICAL CHANGE AND CHEMICAL CHANGE

- No chemical reaction happens.
- No new chemical is formed.

Examples

- Boiling water from the evaporating dish.
- Melting of ice to give water.
- Melting of wax.
- crushing a paper cup.
- crystallisation.

- A chemical reaction happens.
- shape, size, colour, etc may also change

Examples

- Burning of paper.
- Fermentation of grapes.
- souring of milk.
- curd from milk is left in a room during summer.
- Rusting of iron.
- food digestion.
- food cooking.
- chemical Battery usage.

Fast Reactions Primary Processes In Ch

RC Schank



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