

Name Key

1. Hydrogen, H<sub>2</sub>, is used as rocket fuel. The hydrogen is burned in oxygen, O<sub>2</sub>, according to the following equation. What is the enthalpy change if 200.0 g of hydrogen, H<sub>2</sub> is burned?



$$\frac{200.0 \text{ g H}_2}{2.016 \text{ g H}_2} \times \frac{1 \text{ mol H}_2}{1} = 99.2 \text{ mol H}_2$$

$$99.2 \text{ mol H}_2 \times \frac{-484 \text{ kJ}}{2 \text{ mol H}_2} = -24,006 \text{ kJ}$$

-24,006 kJ

2. Carbon disulfide burns in air, producing carbon dioxide and sulfur dioxide.



What is  $\Delta H$  for the following reaction?

reverse so change sign of  $\Delta H$  - it's out in 1/2 quantity



3. Following is the equation for the combustion of acetone:



How much heat will be given off (in kcal) if 4.5 mol of acetone is burned completely?

$$\frac{-853.6 \text{ kcal}}{2 \text{ mol C}_3\text{H}_6\text{O}} \times 4.5 \text{ mol C}_3\text{H}_6\text{O} = -1920.6 \text{ kcal}$$

Exothermic

4. The specific heat of iron is 0.110 cal / g · °C. How much heat must be added to 5.2 g of iron at 25 °C to raise its temperature to 225 °C?

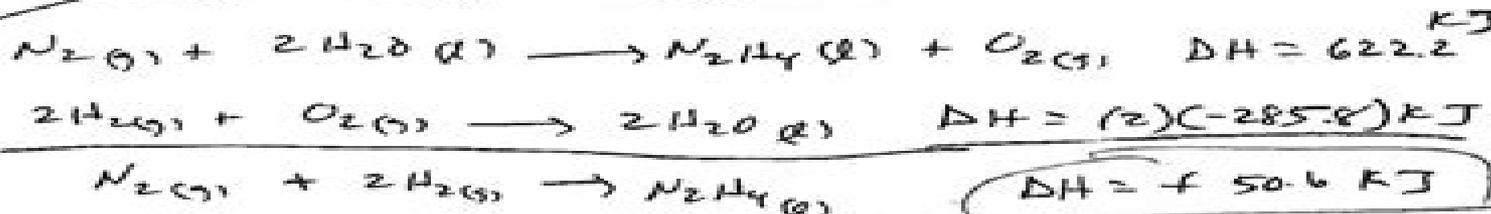
$$\text{heat} = q = m c \Delta T = (5.2 \text{ g}) \left( \frac{0.110 \text{ cal}}{\text{g} \cdot \text{C}} \right) (225 - 25 \text{ C}) = 114.4 \text{ cal}$$

Endothermic

5. Hydrazine, N<sub>2</sub>H<sub>4</sub>, is a colorless liquid used as a rocket fuel. What is the enthalpy change for the process in which hydrazine is formed from its elements (reaction (1) below)? Use Hess's Law to calculate  $\Delta H$  for reaction (1) given the data below.



Reverse  
double quantity



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**Judd E. Hollander**



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## **Table of Contents 19 3 Chemistry Practice Problems Answer Key**

1. Understanding the eBook 19 3 Chemistry Practice Problems Answer Key
  - The Rise of Digital Reading 19 3 Chemistry Practice Problems Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying 19 3 Chemistry Practice Problems Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an 19 3 Chemistry Practice Problems Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from 19 3 Chemistry Practice Problems Answer Key
  - Personalized Recommendations
  - 19 3 Chemistry Practice Problems Answer Key User Reviews and Ratings
  - 19 3 Chemistry Practice Problems Answer Key and Bestseller Lists
5. Accessing 19 3 Chemistry Practice Problems Answer Key Free and Paid eBooks
  - 19 3 Chemistry Practice Problems Answer Key Public Domain eBooks
  - 19 3 Chemistry Practice Problems Answer Key eBook Subscription Services
  - 19 3 Chemistry Practice Problems Answer Key Budget-Friendly Options
6. Navigating 19 3 Chemistry Practice Problems Answer Key eBook Formats

- ePub, PDF, MOBI, and More
  - 19 3 Chemistry Practice Problems Answer Key Compatibility with Devices
  - 19 3 Chemistry Practice Problems Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of 19 3 Chemistry Practice Problems Answer Key
    - Highlighting and Note-Taking 19 3 Chemistry Practice Problems Answer Key
    - Interactive Elements 19 3 Chemistry Practice Problems Answer Key
  8. Staying Engaged with 19 3 Chemistry Practice Problems Answer Key
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers 19 3 Chemistry Practice Problems Answer Key
  9. Balancing eBooks and Physical Books 19 3 Chemistry Practice Problems Answer Key
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection 19 3 Chemistry Practice Problems Answer Key
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine 19 3 Chemistry Practice Problems Answer Key
    - Setting Reading Goals 19 3 Chemistry Practice Problems Answer Key
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of 19 3 Chemistry Practice Problems Answer Key
    - Fact-Checking eBook Content of 19 3 Chemistry Practice Problems Answer Key
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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