

Name Key

1. Hydrogen,  $H_2$ , is used as rocket fuel. The hydrogen is burned in oxygen,  $O_2$ , according to the following equation. What is the enthalpy change if 200.0 g of hydrogen,  $H_2$  is burned?



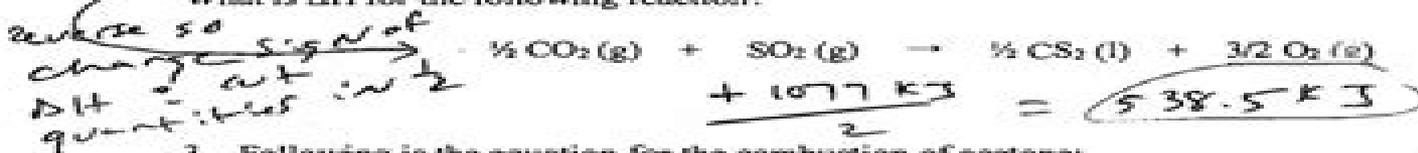
$$\frac{200.0 \text{ g } H_2}{2.016 \text{ g } H_2} \times \frac{1 \text{ mol } H_2}{1} = 99.2 \text{ mol } H_2 \quad \frac{99.2 \text{ mol } H_2}{2 \text{ mol } H_2} \times \frac{-484 \text{ kJ}}{1} = -24,006 \text{ kJ}$$

-24,006 kJ  
~~24,006 kJ~~

2. Carbon disulfide burns in air, producing carbon dioxide and sulfur dioxide.



What is  $\Delta H$  for the following reaction?



3. Following is the equation for the combustion of acetone:



How much heat will be given off (in kcal) if 4.5 mol of acetone is burned completely?

$$\frac{-853.6 \text{ kcal}}{2 \text{ mol } C_3H_6O} \times 4.5 \text{ mol } C_3H_6O = -1920.6 \text{ kcal}$$

Exothermic

4. The specific heat of iron is  $0.110 \text{ cal} / \text{g} \cdot ^\circ\text{C}$ . How much heat must be added to 5.2 g of iron at  $25^\circ\text{C}$  to raise its temperature to  $225^\circ\text{C}$ ?

$$\text{heat} = q = m c \Delta T = (5.2 \text{ g}) \left( \frac{0.110 \text{ cal}}{\text{g} \cdot ^\circ\text{C}} \right) (225 - 25^\circ\text{C}) = 114.4 \text{ cal}$$

Endothermic

5. Hydrazine,  $N_2H_4$ , is a colorless liquid used as a rocket fuel. What is the enthalpy change for the process in which hydrazine is formed from its elements (reaction (1) below)? Use Hess's Law to calculate  $\Delta H$  for reaction (1) given the data below.



# 19 3 Chemistry Practice Problems Answer Key

**RJ Shavelson**



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## **Table of Contents 19 3 Chemistry Practice Problems Answer Key**

1. Understanding the eBook 19 3 Chemistry Practice Problems Answer Key
  - The Rise of Digital Reading 19 3 Chemistry Practice Problems Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying 19 3 Chemistry Practice Problems Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an 19 3 Chemistry Practice Problems Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from 19 3 Chemistry Practice Problems Answer Key

- Personalized Recommendations
  - 19 3 Chemistry Practice Problems Answer Key User Reviews and Ratings
  - 19 3 Chemistry Practice Problems Answer Key and Bestseller Lists
5. Accessing 19 3 Chemistry Practice Problems Answer Key Free and Paid eBooks
- 19 3 Chemistry Practice Problems Answer Key Public Domain eBooks
  - 19 3 Chemistry Practice Problems Answer Key eBook Subscription Services
  - 19 3 Chemistry Practice Problems Answer Key Budget-Friendly Options
6. Navigating 19 3 Chemistry Practice Problems Answer Key eBook Formats
- ePub, PDF, MOBI, and More
  - 19 3 Chemistry Practice Problems Answer Key Compatibility with Devices
  - 19 3 Chemistry Practice Problems Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of 19 3 Chemistry Practice Problems Answer Key
  - Highlighting and Note-Taking 19 3 Chemistry Practice Problems Answer Key
  - Interactive Elements 19 3 Chemistry Practice Problems Answer Key
8. Staying Engaged with 19 3 Chemistry Practice Problems Answer Key
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers 19 3 Chemistry Practice Problems Answer Key
9. Balancing eBooks and Physical Books 19 3 Chemistry Practice Problems Answer Key
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection 19 3 Chemistry Practice Problems Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine 19 3 Chemistry Practice Problems Answer Key
- Setting Reading Goals 19 3 Chemistry Practice Problems Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 19 3 Chemistry Practice Problems Answer Key

- Fact-Checking eBook Content of 19 3 Chemistry Practice Problems Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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