

18.4 Molecular Structure: Mass

Multiple Choice:

1. The molar masses of quantities in chemical equations are based on _____
 (A) percent composition (B) percent yield
 (C) molar formula (D) molar structural analysis
2. Percent of hydrogen in $\text{C}_2\text{H}_6\text{O}$ is _____ of the structural formula.

(A) 15.1% (B) 12.5%

(C) 15.1% (D) 12.5% (E) 15.1% (F) percent of the atoms

3. In a gas the other molecules represent oxygen instead of nitrogen, 2.5 g of gas was produced by the reaction, and only produced 1.5 g of gas. What is the percent yield of sulfur in this reaction?

(A) 60% (B) 60% (C) 60% (D) 60% (E) 60%

(F) 60% (G) 60% (H) 60% (I) 60% (J) 60%

4. In this reaction $\text{ZnCO}_3 \rightarrow \text{ZnO} + \text{CO}_2$, what is the percent yield of oxygen?
 (A) 100% (B) 100% (C) 100% (D) 100%

(E) 100% (F) 100% (G) 100% (H) 100%

5. What is the percent yield of oxygen in this following equation?



(A) 100% (B) 100% (C) 100% (D) 100%

6. Aluminum reacts with HCl to produce hydrogen gas. If 1.5 g of aluminum reacts with HCl, what mass of H₂ gas is produced?



(A) 1.5 g (B) 1.5 g (C) 1.5 g (D) 1.5 g (E) 1.5 g

7. In a full experiment, structural formula of $\text{C}_2\text{H}_6\text{O}$ is shown. What is the percent yield of oxygen in this reaction? (A) 100% (B) 100% (C) 100% (D) 100%



Multiple Choice Stoichiometry Test

M Walker



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