

## 19.2

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

- 1. Calculate  $K_a$  and  $K_b$  from  $K_w$  and  $K_{sp}$  for weak acids and bases.
- 2. Calculate the pH of weak acids and bases.
- 3. Calculate the percent ionization.

## Vocabulary

- |                 |                   |      |
|-----------------|-------------------|------|
| acid            | 1. conjugate base | acid |
| acidic solution | 2. conjugate acid | base |
| base            | 3. indicator      |      |

## Key Equations

- |                                    |                                   |
|------------------------------------|-----------------------------------|
| 1. $K_a = \frac{[H^+][A^-]}{[HA]}$ | 2. $K_b = \frac{[HB][OH^-]}{[B]}$ |
| 3. $K_w = [H^+][OH^-]$             | 4. $K_a \cdot K_b = K_w$          |

## Part A Completion

Write the conjugate acids to the following bases. Write the conjugate bases to the following acids. Write the  $K_a$  and  $K_b$  expressions for the following acids and bases.

1.  $SO_4^{2-}$
2.  $CO_3^{2-}$
3.  $HSO_4^-$
4.  $H_2PO_4^-$
5.  $H_2O$
6.  $H_2O$
7.  $H_2O$
8.  $H_2O$
9.  $H_2O$
10.  $H_2O$
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39.  $H_2O$
40.  $H_2O$

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**Michael Seilmaier**



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## **Table of Contents 19 2 Hydrogen Ions And Acidity Pearson**

1. Understanding the eBook 19 2 Hydrogen Ions And Acidity Pearson
  - The Rise of Digital Reading 19 2 Hydrogen Ions And Acidity Pearson
  - Advantages of eBooks Over Traditional Books
2. Identifying 19 2 Hydrogen Ions And Acidity Pearson
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an 19 2 Hydrogen Ions And Acidity Pearson
  - User-Friendly Interface
4. Exploring eBook Recommendations from 19 2 Hydrogen Ions And Acidity Pearson
  - Personalized Recommendations
  - 19 2 Hydrogen Ions And Acidity Pearson User Reviews and Ratings
  - 19 2 Hydrogen Ions And Acidity Pearson and Bestseller Lists
5. Accessing 19 2 Hydrogen Ions And Acidity Pearson Free and Paid eBooks
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  - 19 2 Hydrogen Ions And Acidity Pearson eBook Subscription Services
  - 19 2 Hydrogen Ions And Acidity Pearson Budget-Friendly Options
6. Navigating 19 2 Hydrogen Ions And Acidity Pearson eBook Formats

- ePub, PDF, MOBI, and More
  - 19 2 Hydrogen Ions And Acidity Pearson Compatibility with Devices
  - 19 2 Hydrogen Ions And Acidity Pearson Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of 19 2 Hydrogen Ions And Acidity Pearson
    - Highlighting and Note-Taking 19 2 Hydrogen Ions And Acidity Pearson
    - Interactive Elements 19 2 Hydrogen Ions And Acidity Pearson
  8. Staying Engaged with 19 2 Hydrogen Ions And Acidity Pearson
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers 19 2 Hydrogen Ions And Acidity Pearson
  9. Balancing eBooks and Physical Books 19 2 Hydrogen Ions And Acidity Pearson
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection 19 2 Hydrogen Ions And Acidity Pearson
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine 19 2 Hydrogen Ions And Acidity Pearson
    - Setting Reading Goals 19 2 Hydrogen Ions And Acidity Pearson
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of 19 2 Hydrogen Ions And Acidity Pearson
    - Fact-Checking eBook Content of 19 2 Hydrogen Ions And Acidity Pearson
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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