



# The Banished Princess

**Hussin A.Rothana**



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web elements in the second row of the periodic table place their electrons in the  $2n$  shell as well as the  $1n$  shell for instance lithium li has three electrons two fill the  $1s$  orbital and the third is placed in the  $2s$  orbital giving an electron configuration of  $1s^2 2s^1$

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web one more point needs to be emphasized about the relationship between electron configuration and the periodic table the atoms of elements in the same vertical column of the table have similar electron configurations for example consider the alkaline earth elements group iia

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web the electron removed during the ionization of beryllium  $he\ 2s^2$  is an  $s$  electron whereas the electron removed during the ionization of boron  $he\ 2s^2 2p^1$  is a  $p$  electron this results in a lower first ionization energy for boron even though its nuclear charge is greater by one proton

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web predictions from reliable sources have been used for these elements grayed out electron numbers indicate subshells filled to their maximum bracketed noble gas symbols on the left represent inner configurations that are the same in each period written out these are  $he\ 2$  helium  $1s^2$   $ne\ 10$  neon  $1s^2 2s^2 2p^6$

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web q1 what is the electron configuration of the element potassium answer the atomic number of potassium is 19 a neutral potassium element has 19 electrons its electron configuration is  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1$  or simply put  $ar\ 4s^1$  q2 what is the total number of electrons that can be filled in  $s$   $p$  and  $d$  subshells

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web by building up from hydrogen this table can be used to determine the electron configuration for any atom on the periodic table we will now construct the ground state electron configuration and orbital diagram for a selection of atoms in the first and second periods of the periodic table

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web electron configurations are a simple way of writing down the locations of all of the electrons in an atom as we know the positively charged protons in the nucleus of an atom tend to attract negatively charged electrons

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