

Voltaic Cells

Label the parts for the following cell:



By connecting to the voltmeter, describe a spontaneous redox reaction. (10 marks)

- a. metals
- b. solutions
- c. electrodes
- d. voltmeter
- e. salt bridge
- f. zinc sulfate solution
- g. copper sulfate solution

g. zinc sulfate solution

h. Cu

i. Both metals and their solutions are placed inside separate beakers connected by a salt bridge which will maintain the original charge.

Voltaic Cells Study Guide

PT Brinkman



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Voltaic Cells Study Guide

web q5 identify the reducing agent in the following reactions
a $4 \text{NH}_3 + 5 \text{O}_2 \rightarrow 4 \text{NO} + 6 \text{H}_2\text{O}$ b $2 \text{H}_2\text{O} + \text{HF} \rightarrow \text{H}_3\text{O}^+ + \text{F}^-$ c $2 \text{O}_3 \rightarrow 3 \text{O}_2$ d $2 \text{H}_2\text{O} \rightarrow 2 \text{H}_2 + \text{O}_2$
answer a here ammonia NH_3 is the reducing agent b here water H_2O is the reducing agent