

The number of moles of substances, when compared and reduced to the lowest terms, yields a mole ratio, as shown in the sample below. Indicate the mole ratio for each of the following equations:



$$\underline{2:1:2}$$



$$\underline{1:1:1}$$



$$\underline{4:3:2}$$



$$\underline{1:2:1:1}$$



$$\underline{2:1:2}$$



$$\underline{2:1:1}$$

Convert the following masses to moles by using the information on your periodic table.

a) 34.0 g NH_3 FW = 17.03/mol

$$34.0\text{g} \left(\frac{1\text{mol}}{17.0\text{g}} \right) = 2.00\text{mol}$$

$$\underline{2.00\text{mol}}$$

b) 51.4 g Na_2O FW = 62.0g/mol

$$51.4\text{g} \left(\frac{1\text{mol}}{62.0\text{g}} \right) = 0.829\text{mol}$$

$$\underline{0.829\text{mol}}$$

c) 4.50 g H_2SO_4 FW = 98.1 g/mol

$$4.50\text{g} \left(\frac{1\text{mol}}{98.1\text{g}} \right) = 0.0459\text{mol}$$

$$\underline{0.0459\text{mol}}$$

d) 229 g $\text{Ca}(\text{NO}_3)_2$ FW = 164.1g/mol

$$229\text{g} \left(\frac{1\text{mol}}{164.1\text{g}} \right) = 1.40\text{mol}$$

$$\underline{1.40\text{mol}}$$

e) 79.5 g NH_4OH FW = 35.0g/mol

$$79.5\text{g} \left(\frac{1\text{mol}}{35.0\text{g}} \right) = 2.27\text{mol}$$

$$\underline{2.27\text{mol}}$$

Convert the following to grams.

a) 17.1 moles H_2O FW = 18.0g/mol

$$17.1\text{mol} \times \left(\frac{18.0\text{g}}{1\text{mol}} \right) = 308\text{g}$$

$$\underline{308\text{g}}$$

Section 3 1 Chemistry Answers

Henry Aubrey Husband



Section 3 1 Chemistry Answers:

Unveiling the Energy of Verbal Artistry: An Mental Sojourn through **Section 3 1 Chemistry Answers**

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