

16.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations and pH values and vice versa for weak acids and bases.
3. Calculate the pH of weak acid solutions.

Vocabulary

- | | | |
|-------------------|------------------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid |
| 4. hydronium | 5. hydroxide ion (OH^-) | 6. autoionization |
| | 7. weak base | 8. pH |

Key Equations

- | | |
|----------------------------------------------------------------------------------------------------|----------------------------------------------------------------------------------------------------|
| 1. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ | 2. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ |
| 3. $\text{pH} = -\log[\text{H}_3\text{O}^+]$ | 4. $\text{pH} = -\log[\text{H}_3\text{O}^+]$ |

Part A: Completion

Use the appropriate number of significant figures in the answers to the following questions that are based on the following acid-base reactions:

1. What products are **1** and **2** in the following hydrolysis reaction?
 $\text{HCO}_3^- + \text{H}_2\text{O} \rightleftharpoons \text{H}_2\text{CO}_3 + \text{OH}^-$
2. What products are **1** and **2** in the autoionization of water which produces an H_3O^+ ion and an OH^- ion?
 $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$
3. What products are **1** and **2** in the following reaction?
 $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$
4. What products are **1** and **2** in the following reaction?
 $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$
5. What products are **1** and **2** in the following reaction?
 $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$
6. What products are **1** and **2** in the following reaction?
 $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$

1. H_2CO_3
2. OH^-
3. H_3O^+
4. OH^-
5. H_2O
6. H_2O
7. H_3O^+
8. OH^-
9. H_3O^+
10. OH^-
11. H_3O^+
12. OH^-

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