

18.4 Acid-Base Chemistry: Theory

Multiple-Choice Questions

1. Which combination of quantities are dimensional quantities? (correct)
- (A) pressure, concentration, pH, amount of substance
 (B) volume, molarity, pH, amount of substance
 (C) volume, molarity, pH, concentration of ions
 (D) volume, molarity, pH, concentration of ions, amount of substance

(A) $V = \text{dm}^3$, $M = \text{mol dm}^{-3}$

(B) $V = \text{dm}^3$, $n = \text{mol}$, $pH = -\log_{10} \frac{[H^+]}{1 \text{ mol dm}^{-3}}$, $c = \text{mol dm}^{-3}$, $[H^+] = 10^{-pH} \text{ mol dm}^{-3}$
 (C) $V = \text{dm}^3$, $n = \text{mol}$, $pH = -\log_{10} \frac{[H^+]}{1 \text{ mol dm}^{-3}}$, $c = \text{mol dm}^{-3}$, $[H^+] = 10^{-pH} \text{ mol dm}^{-3}$

2. An acid-base indicator has a pK_a of 4.5. In a solution of the indicator, the ratio of the concentrations of the two forms is 10:1. What is the pH of the solution? (correct)

(A) $pH = 4.5 + \log_{10} 10$ (B) $pH = 4.5 - \log_{10} 10$

(C) $pH = 4.5$ (D) $pH = 4.5 + \log_{10} 10$ (E) $pH = 4.5 - \log_{10} 10$

3. An acid-base indicator has a pK_a of 4.5. In a solution of the indicator, the ratio of the concentrations of the two forms is 10:1. What is the pH of the solution? (correct)

(A) 4.5 (B) 5.5 (C) 3.5 (D) 5.0

4. Which indicator is most suitable for the titration of a weak acid with a strong base? (correct)

(A) $pK_a = 3.5$ (B) $pK_a = 7.5$

(C) 3 (D) 7

5. A solution of a weak acid and its conjugate base is prepared by mixing 100 cm³ of 0.100 mol dm⁻³ of the weak acid with 100 cm³ of 0.100 mol dm⁻³ of the conjugate base. The pK_a of the weak acid is 4.5. What is the pH of the solution? (correct)

(A) 4.5 (B) 5.5 (C) 3.5 (D) 7.5

(E) 4.5 + 0.5 (F) 4.5 - 0.5 (G) 3.5 + 0.5 (H) 7.5 - 0.5

6. An acid-base equilibrium is shown below. What is the pK_a of the acid? (correct)



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