

Candidates are required to give their answers in their own words as far as practicable. The figures in the margin indicate full marks.

Attempt (Any Five) Questions.

1. a) Define the terms; Electric current, voltage, power and energy. [8]
b) Define Resistor and capacitor. What will be the current drawn by a lamp rated at 250v, 40 watts connected to a 230v supply? [4+4]
2. a) State and prove Ohm's law with suitable example. [8]
b) The effective resistance of two resistors connected in series is 100Ω . When connected in parallel the effective value is 24Ω . Determine the values of the two resistors. [8]
3. a) Differentiate between Dia, para and ferro-magnetic material. [8]
b) Derive series and parallel connection of capacitor. [8]
4. a) Define magnetic field and magnetic flux. Also explain the hysteresis loop for magnetic material. [8]
b) Derive the energy stored in current carrying inductor. [8]
5. a) What is Lenz's law? Explain dynamically induced emf and statistically induced emf. [10]
b) Explain in detail about Faraday's laws of Electromagnetic induction. [6]
6. Write short notes on : (Any Four) [4x4=16]
 - a) resistance and its variation with temperature
 - b) Electrostatic induction phenomena.
 - c) KVL circuit
 - d) Faraday's law of electrolysis
 - e) Lead acid cell
 - f) Soft and hard magnetic material

Good luck !

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H Kauffman



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Table of Contents Diploma In Electrical Question Paper

1. Understanding the eBook Diploma In Electrical Question Paper
 - The Rise of Digital Reading Diploma In Electrical Question Paper
 - Advantages of eBooks Over Traditional Books
2. Identifying Diploma In Electrical Question Paper
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Diploma In Electrical Question Paper
 - User-Friendly Interface
4. Exploring eBook Recommendations from Diploma In Electrical Question Paper

- Personalized Recommendations
 - Diploma In Electrical Question Paper User Reviews and Ratings
 - Diploma In Electrical Question Paper and Bestseller Lists
5. Accessing Diploma In Electrical Question Paper Free and Paid eBooks
 - Diploma In Electrical Question Paper Public Domain eBooks
 - Diploma In Electrical Question Paper eBook Subscription Services
 - Diploma In Electrical Question Paper Budget-Friendly Options
 6. Navigating Diploma In Electrical Question Paper eBook Formats
 - ePub, PDF, MOBI, and More
 - Diploma In Electrical Question Paper Compatibility with Devices
 - Diploma In Electrical Question Paper Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Diploma In Electrical Question Paper
 - Highlighting and Note-Taking Diploma In Electrical Question Paper
 - Interactive Elements Diploma In Electrical Question Paper
 8. Staying Engaged with Diploma In Electrical Question Paper
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Diploma In Electrical Question Paper
 9. Balancing eBooks and Physical Books Diploma In Electrical Question Paper
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Diploma In Electrical Question Paper
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Diploma In Electrical Question Paper
 - Setting Reading Goals Diploma In Electrical Question Paper
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Diploma In Electrical Question Paper

- Fact-Checking eBook Content of Diploma In Electrical Question Paper
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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web with a strong acid and strong base titration strong acids and strong bases refer to species that completely dissociate to form ions in solution by contrast weak acids and bases ionize only partially and the ionization reaction is reversible

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web conclusion the purpose of the experiment was to identify the unknown solid acid 17 by calculating its molar mass this was done by titration of the acid with sodium hydroxide in order to find the volume of sodium hydroxide required to reach the end point and that value was used in unit analysis to find the molar mass and identify the unknown solid acid

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web by utilizing the half equivalence point on the titration curve the pka of the acid analyte was determined to be 3.655 furthermore using the stoichiometric ratio between the naoh and the acid analyte 1:1 we could experimentally determine the number of moles of the acid in the solution to be 0.00731 moles making the experimental molar mass

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web goals understand that parts of an acid base titration be able to determine the K_a or K_b from pH data associated with the titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to calculate K_{a1} and K_{a2} for a polyprotic acid by the end of this lab students should be able to

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web titration experiment in the neutralization of hydrochloric acid by sodium hydroxide the mole ratio of acid to base is 1:1 $HCl(aq) + NaOH(aq) \rightarrow NaCl(aq) + H_2O(l)$ one mole of HCl would be fully neutralized by one mole of NaOH

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web conclusion the identification of an unknown solid acid the purpose of this experiment was to determine the standardization of sodium hydroxide NaOH and oxalic acid $H_2C_2O_4 \cdot 2H_2O$ and the molar mass of an unknown solid acid using the molar mass of the standardized sodium hydroxide solution

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web acid base reactions can be used to determine concentrations of an acid or a base using stoichiometric calculations and titrimetric analysis the volume of a solution required to react

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web specifically an acid base titration can be used to figure out the concentration of an acid or base whether an unknown acid or base is strong or weak pK_a of an unknown acid or pK_b of the unknown base key terms titration a process where a solution of known strength is added to a certain volume of a treated sample containing an

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web conclusion this lab evaluated an acid base reaction with the volumetric analysis of a standard solution being naoh reaction with a 2nd substance $\text{H}_2\text{C}_2\text{O}_4 \text{ aq} + 2\text{NaOH} \text{ aq} \rightarrow \text{Na}_2\text{C}_2\text{O}_4 + 4\text{H}_2\text{O}$ to differentiate an acid from a base in a solution a ph scale is used the ph scale ranges from 0 to 14 a ph of 7 is neutral in an acid base reaction

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web this point of titration is referred to as the equivalence point considering the 1 1 stoichiometry of this acid base reaction $\text{NaOH} \text{ aq} + \text{C}_6\text{H}_4\text{COOH} \text{ aq} \rightarrow \text{C}_6\text{H}_4\text{COONa} \text{ aq} + \text{H}_2\text{O} \text{ l}$ the point of equivalence is the point of titration when the number of moles of naoh na added is equal to the number of moles of khp nb in the