

Procedure #	Liquid or Solid Used	Reagent	Observation	Ppt contains	Centrifugate Contains	Applicable Equations
1	Group I, II, III, IV	HCl	White & Yellow	AgCl, Hg ₂ Cl ₂ , PbCl ₂	Group 2, 3, and 4	$Ag^+ + HCl \rightarrow AgCl + H^+$ $2Hg^{2+} + 2HCl \rightarrow Hg_2Cl_2 + 2H^+$ $Pb^{2+} + 2HCl \rightarrow PbCl_2 + 2H^+$
2	PPT From #1	Hot H ₂ O	PPT Still Present	AgCl, Hg ₂ Cl ₂	Pb ²⁺	$PbCl_2 + H_2O + Heat \rightarrow Pb^{2+}(aq) + Cl^-(aq)$
3	Half of decant from #2	K ₂ CrO ₄	Yellow Precipitate	PbCrO ₄	N/A	$Pb^{2+}(aq) + K_2CrO_4 \rightarrow PbCrO_4 + 2K^+$
	Other half of decant	H ₂ SO ₄	White Precipitate	PbSO ₄	N/A	$H_2SO_4 + Pb^{2+} \rightarrow PbSO_4 + 2H^+$
4a	PPT From #2	NH ₃	Black/Gray PPT	Hg ⁺ HgNH ₂ Cl	Ag(NH ₂) ⁺	$AgCl + 2NH_3 \rightarrow Ag(NH_2)^+ + Cl^-$ $Hg_2Cl_2 + NH_3 \rightarrow Hg(l) + HgNH_2Cl(s) + NH_4^+ + Cl^-$
4b	Decantate	HNO ₃	White Precipitate	AgCl	N/A	$Ag(NH_2)^+(aq) + 2H^+(aq) \rightarrow Ag^+(aq) + 2NH_4^+(aq)$ $Ag^+(aq) + Cl^-(aq) \rightarrow AgCl(s)$
5a	Decant from #1	H ₂ O ₂ , (NH ₄) ₂ S, NH ₄ C ₂ H ₃ O ₂ , H ₂ O, Heat	Dark/Hunter Green	PbS, Bi ₂ S ₃ , CuS, CdS	Group III and IV	$Pb^{2+} + S^{2-} \rightarrow PbS$ $2Bi^{3+} + 3S^{2-} \rightarrow Bi_2S_3$ $Cu^{2+} + S^{2-} \rightarrow CuS$ $Cd^{2+} + S^{2-} \rightarrow CdS$
5b						
6a	PPT From #5	HNO ₃ , H ₂ SO ₄ , Heat	Yellow Crystal Forms, ions dissolve, and White Smoke Appears	Sulfur	Group II ions	$PbS + 2NO_3^- + 4H^+ \rightarrow Pb^{2+} + S + 2NO_2 + 2H_2O$ $CuS + 2NO_3^- + 4H^+ \rightarrow Cu^{2+} + S + 2NO_2 + 2H_2O$ $CdS + 2NO_3^- + 4H^+ \rightarrow Cd^{2+} + S + 2NO_2 + 2H_2O$ $Bi_2S_3 + 6NO_3^- + 12H^+ \rightarrow 2Bi^{3+} + 3S + 6NO_2 + 6H_2O$
6b	Decant from 6a	Ice, NH ₃ , HC ₂ H ₃ O ₂ , KCrO ₄	White Precipitate, Then Dissolve, Then Yellow PPT	PbSO ₄	Bi ³⁺ , Cu ²⁺ , Cd ²⁺	$Pb^{2+} + SO_4^{2-} + ICE \rightarrow PbSO_4 + Heat$ $PbSO_4 + 3NH_3 + C_2H_3O_2 \rightarrow Pb(C_2H_3O_2)_2 + 3NH_4^+ + SO_4^{2-}$ $Pb(C_2H_3O_2)_2 + KCrO_4 \rightarrow PbCrO_4 + 3[C_2H_3O_2]^- + K^+$
7a	Decant from #6	NH ₃				$Bi^3 + 3NH_3 + 3H_2O \rightarrow Bi(OH)_3 + 3NH_4^+$
		NaOH, SnCl ₂	Blue	Bi(OH) ₃ (s)	Cu(NH ₂) ₂ ⁺ , Cd(NH ₂) ₂ ⁺	$Cu^{2+} + 4NH_3 \rightarrow Cu(NH_2)_2^{2+}$ (Blue) $Cd^{2+} + 4NH_3 \rightarrow Cd(NH_2)_2^{2+}$
7b	PPT From #7a	H ₂ O, NaOH, SnCl ₂	Jet Black	Bi(S)	N/A	$SnCl_2 + 3NaOH \rightarrow Sn(OH)_2 + 3Na^+ + 2Cl^-$ $BiOH + SnOH \rightarrow Sn(OH)_2^{2-} + Bi^0(S)$

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