

18.4 Molecular Structure: Mass

Multiple Choice:

1. The molar masses of quantities in chemical equations are based _____
 (A) percent composition (B) percent yield
 (C) molar mass (D) molar volume
2. Percent of oxygen in $\text{CaSO}_4 \cdot \frac{1}{2}\text{H}_2\text{O}$ is approximately (to two significant figures) _____
 (A) 20% (B) 25% (C) 30% (D) 35%

3. In a gas the other reactions represent copper metal and silver chloride, 2.5 g of Ag are completely reacted by the reaction, and only 0.5 g of AgCl formed. What mass of copper is present in the reaction?
 (A) 2.0 g (B) 2.5 g (C) 3.0 g (D) 3.5 g

4. In a gas the other reactions represent copper metal and silver chloride, 2.5 g of Ag are completely reacted by the reaction, and only 0.5 g of AgCl formed. What mass of copper is present in the reaction?
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6. In a gas the other reactions represent copper metal and silver chloride, 2.5 g of Ag are completely reacted by the reaction, and only 0.5 g of AgCl formed. What mass of copper is present in the reaction?
 (A) 2.0 g (B) 2.5 g (C) 3.0 g (D) 3.5 g

7. Percent of iron in Fe_2O_3 is approximately (to two significant figures) _____
 (A) 60% (B) 70% (C) 80% (D) 90%

8. A mixture of 10.0 g of Fe and 10.0 g of S is heated until the reaction is complete. The mass of FeS formed is _____
 (A) 10.0 g (B) 20.0 g (C) 30.0 g (D) 40.0 g

9. In a gas the other reactions represent copper metal and silver chloride, 2.5 g of Ag are completely reacted by the reaction, and only 0.5 g of AgCl formed. What mass of copper is present in the reaction?
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Multiple Choice Stoichiometry Test

Daniela Niemeyer



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