

Answer Sheet for Sample Problems for Chemistry Level 1 Final Exam 2013-2014 Study Guide

Electrons in Atoms – Chapter 5

1. What is the frequency of green light, which has a wavelength of $4.90 \times 10^{-7} \text{ m}$?

$$c = \lambda \nu \quad \lambda = 4.90 \times 10^{-7} \text{ m}, c = 3.00 \times 10^8 \text{ m/s} \quad \nu = ? \text{ Hz} \quad \nu = \frac{c}{\lambda} = \frac{3.00 \times 10^8 \text{ m/s}}{4.90 \times 10^{-7} \text{ m}}$$

$$\underline{6.12 \times 10^{14} \text{ Hz}}$$

2. An x-ray has a wavelength of $1.15 \times 10^{-10} \text{ m}$. What is its frequency?

$$c = \lambda \nu \quad \lambda = 1.15 \times 10^{-10} \text{ m}, c = 3.00 \times 10^8 \text{ m/s} \quad \nu = ? \text{ Hz} \quad \nu = \frac{c}{\lambda} = \frac{3.00 \times 10^8 \text{ m/s}}{1.15 \times 10^{-10} \text{ m}}$$

$$\underline{2.61 \times 10^{18} \text{ Hz}}$$

3. Which type of electromagnetic radiation has the most energy: Microwaves, X-Rays, or Visible light? Why?

Ans: X-Rays, b/c they have the shortest wavelength and highest frequency of the types listed and frequency and energy are directly related.

4. What is the energy of a $9.50 \times 10^{13} \text{ Hz}$ radiation?

$$E = h\nu \quad \nu = 9.50 \times 10^{13} \text{ Hz}, h = 6.626 \times 10^{-34} \text{ J s} \quad E = ? \text{ J}$$

$$E = h\nu = (6.626 \times 10^{-34} \text{ J s}) (9.50 \times 10^{13} \text{ s}^{-1}) = \underline{6.29 \times 10^{-19} \text{ J}}$$

5. The element with the electron configuration of $[\text{Kr}]5s^24d^{10}5p^1$ is indium (In)

6. Barium (Ba)

7. Write out the electron configurations for the following elements: (For the exam make sure you know how to do both the complete and the noble gas short cut for electron configurations)

a. Tc **Ans: $1s^22s^22p^63s^23p^44s^23d^54p^65s^24d^5$ or $[\text{Kr}]5s^24d^5$**

b. Aluminum **Ans: $1s^22s^22p^63s^23p^1$ or $[\text{Ne}] 3s^23p^1$**

8. Why are Cr and Cu exceptions to writing the electron configurations of elements?

They are exceptions because instead of following the pattern of filling up the d orbitals Cr actually has a half filled s and d orbitals and Cu has a half filled s and completely filled d orbitals. The reason why this occurs is because the electrons in the d-sublevel are more stable with this electron configuration than the expected configuration.

9. a. Ca

b. As

c. Sn

a. 2 Ca.

b. 5 .As.

c. 4 . Sn .

Periodic Table and Periodic Trends – Chapter 6

1. Compare & contrast how Mendeleev's periodic table was organized to how Moseley organized it.

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Table of Contents 2nd Semester Exam Review Sheet Answer

1. Understanding the eBook 2nd Semester Exam Review Sheet Answer
 - The Rise of Digital Reading 2nd Semester Exam Review Sheet Answer
 - Advantages of eBooks Over Traditional Books
2. Identifying 2nd Semester Exam Review Sheet Answer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 2nd Semester Exam Review Sheet Answer
 - User-Friendly Interface

4. Exploring eBook Recommendations from 2nd Semester Exam Review Sheet Answer
 - Personalized Recommendations
 - 2nd Semester Exam Review Sheet Answer User Reviews and Ratings
 - 2nd Semester Exam Review Sheet Answer and Bestseller Lists
5. Accessing 2nd Semester Exam Review Sheet Answer Free and Paid eBooks
 - 2nd Semester Exam Review Sheet Answer Public Domain eBooks
 - 2nd Semester Exam Review Sheet Answer eBook Subscription Services
 - 2nd Semester Exam Review Sheet Answer Budget-Friendly Options
6. Navigating 2nd Semester Exam Review Sheet Answer eBook Formats
 - ePub, PDF, MOBI, and More
 - 2nd Semester Exam Review Sheet Answer Compatibility with Devices
 - 2nd Semester Exam Review Sheet Answer Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 2nd Semester Exam Review Sheet Answer
 - Highlighting and Note-Taking 2nd Semester Exam Review Sheet Answer
 - Interactive Elements 2nd Semester Exam Review Sheet Answer
8. Staying Engaged with 2nd Semester Exam Review Sheet Answer
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 2nd Semester Exam Review Sheet Answer
9. Balancing eBooks and Physical Books 2nd Semester Exam Review Sheet Answer
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 2nd Semester Exam Review Sheet Answer
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 2nd Semester Exam Review Sheet Answer
 - Setting Reading Goals 2nd Semester Exam Review Sheet Answer
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of 2nd Semester Exam Review Sheet Answer
 - Fact-Checking eBook Content of 2nd Semester Exam Review Sheet Answer
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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