



# Hazardous Material Transport

**RC Schank**



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web solution a titrant volume 0 ml the solution pH is due to the acid ionization of HCl because this is a strong acid the ionization is complete and the hydronium ion molarity is 0.100 M the pH of the solution is then  $pH = -\log 0.100 = 1.000$  b titrant volume 12.50 ml

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web expert answer 100% 2 ratings the example is based upon Henderson-Hasselbalch equation the equation gives relation between pH and  $pK_a$  of weak acid the equation is  $pH = pK_a + \log \frac{[salt]}{[acid]}$  at equivalence point all a view the full answer transcribed image text

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