

18.4 Acid-Base Chemistry: Theory

Multiple-Choice Questions

1. Which combination of quantities are dimensional quantities? (correct)
- (A) pressure, concentration, pH, amount of substance
 (B) volume, molarity, pH, amount of substance
 (C) volume, molarity, pH, concentration of ions
 (D) amount of substance, pH, concentration of ions, concentration of substance

(A) $\text{pH} = \frac{[\text{H}^+]}{10^{-7}}$ (B) $\text{pH} = \frac{[\text{H}^+]}{10^{-14}}$

(C) $\text{pH} = \frac{[\text{H}^+]}{10^{-7}}$ (D) $\text{pH} = \frac{[\text{H}^+]}{10^{-14}}$ (E) $\text{pH} = \frac{[\text{H}^+]}{10^{-7}}$ (F) $\text{pH} = \frac{[\text{H}^+]}{10^{-14}}$

2. An acid-base reaction involves hydrogen ions. Suppose a solution with a pH of 1.0 is mixed with a solution with a pH of 2.0. What is the pH of the resulting solution? (correct)

(A) $\text{pH} = 1.5$ (B) $\text{pH} = 1.3$ (C) $\text{pH} = 1.7$ (D) $\text{pH} = 1.9$ (E) $\text{pH} = 1.6$

(F) $\text{pH} = 1.8$ (G) $\text{pH} = 1.4$ (H) $\text{pH} = 1.5$ (I) $\text{pH} = 1.6$ (J) $\text{pH} = 1.7$

3. An acid-base reaction involves hydrogen ions. Suppose a solution with a pH of 1.0 is mixed with a solution with a pH of 2.0. What is the pH of the resulting solution? (correct)

(A) $\text{pH} = 1.5$ (B) $\text{pH} = 1.3$ (C) $\text{pH} = 1.7$ (D) $\text{pH} = 1.9$ (E) $\text{pH} = 1.6$

4. Which of the following is a weak acid? (correct)



5. A solution of a weak acid and its conjugate base is prepared. The pH of the solution is 4.0. What is the ratio of the concentration of the weak acid to the concentration of the conjugate base? (correct)

(A) 10 (B) 100 (C) 1000 (D) 10000 (E) 100000

(F) 1000000 (G) 10000000 (H) 100000000 (I) 1000000000 (J) 10000000000

6. An acid-base reaction involves hydrogen ions. Suppose a solution with a pH of 1.0 is mixed with a solution with a pH of 2.0. What is the pH of the resulting solution? (correct)

(A) $\text{pH} = 1.5$ (B) $\text{pH} = 1.3$ (C) $\text{pH} = 1.7$ (D) $\text{pH} = 1.9$ (E) $\text{pH} = 1.6$

Multiple Choice Stoichiometry Test

Richard C. Sweetland, Daniel J. Keyser

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