

The curve in Figure 16.11 of the text shows how the vibrational heat capacity depends on temperature. Note that even when the temperature is only slightly above the vibrational temperature, the heat capacity is close to its equipartition value.

The total heat capacity of a molecular substance is the sum of each contribution (Figure 16.12 of the text). When equipartition is valid (when the temperature is well above the characteristic temperature of the mode $T \gg \theta_M$) we can estimate the heat capacity by counting the numbers of modes that are active. In gases, all three translational modes are always active and contribute $\frac{3}{2}R$ to the molar heat capacity. If we denote the number of active rotational modes by ν_R^\ddagger (so for most molecules at normal temperatures $\nu_R^\ddagger = 2$ for linear molecules, and 3 for non-linear molecules), then

$$C_{V,m} = \frac{3}{2}R + \frac{\nu_R^\ddagger}{2}R + \dots$$

The equipartition value $C_{V,m} = \frac{3}{2}R + \frac{\nu_R^\ddagger}{2}R$ is reached when the temperature is high enough so that $T \gg \theta_M$ for all the modes. However, if the temperature is lower than the characteristic temperature of some of the modes, then the heat capacity is lower than the equipartition value. The heat capacity is lower than the equipartition value because the energy of the modes is not fully excited. The energy of the modes is not fully excited because the temperature is lower than the characteristic temperature of some of the modes. The energy of the modes is not fully excited because the temperature is lower than the characteristic temperature of some of the modes.

Solutions to exercises

The characteristic temperature is given by $\theta_M = \frac{h\nu}{k_B}$. The characteristic temperature is given by $\theta_M = \frac{h\nu}{k_B}$. The characteristic temperature is given by $\theta_M = \frac{h\nu}{k_B}$.

$$\theta_M = \frac{h\nu}{k_B} = \frac{6.626 \times 10^{-34} \text{ J s} \times 1.0 \times 10^{14} \text{ s}^{-1}}{1.38 \times 10^{-23} \text{ J K}^{-1}} = 4.8 \text{ K}$$

At 298 K, $T \gg \theta_M$, so the equipartition value is reached.

$$C_{V,m} = \frac{3}{2}R + \frac{2}{2}R = \frac{5}{2}R = 20.8 \text{ J K}^{-1} \text{ mol}^{-1}$$

At 100 K, $T \approx \theta_M$.

$$C_{V,m} = \frac{3}{2}R + \frac{2}{2}R \left(\frac{T}{\theta_M} \right)^2 = \frac{5}{2}R \left(\frac{100}{4.8} \right)^2 = 20.8 \text{ J K}^{-1} \text{ mol}^{-1}$$

At 10 K, $T \ll \theta_M$.

$$C_{V,m} = \frac{3}{2}R = 12.5 \text{ J K}^{-1} \text{ mol}^{-1}$$

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Mike Jess



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