

Physical Science – 2nd Semester – Final Exam Study Guide – (2012–2013)**1. Physical Science Basics**

- a. What tool(s) would you want to use to find the:
- Mass of an object? **triple beam balance** Basic SI Unit? **kg**
 - Volume of an object? **ruler or water displacement method** Basic SI Unit? **cm³ or mL**
 - Length of an object? **Ruler** Basic SI Unit? **meter**
 - Density of an object? **graduated cylinder** Basic SI Unit? **mL**
- b. What is the formula for calculating density? **mass/volume**
- What is the density of a cube that is 35 g and has a volume of 3 cm³? **35/3cm³**
 - Compare the density of solid H₂O to liquid H₂O, and explain what observable effect this has. **solid H₂O less dense than liquid H₂O**

2. Chemistry Basics

- a. Element (def.) **substance made up of one kind of atom**
- b. Atom (def.) **basic building block of matter**
- c. Compound (def.) **a molecule that contains atoms or molecules of two or more different elements**
- d. Molecule (def.) **a combination of two or more of the same or different atoms that are bonded together**

3. Model of the Atomic Evolution of the Atomic Theory

Atoms are so small that many scientists throughout history have created **models** to describe them.

Why have atomic theories changed over time? **better technology came around.**

Scientist	Year	Major Discovery/ Contribution	Name of atomic model
John Dalton	1808	Each element is made up of small, indivisible particles called atoms. Different elements have atoms of different masses.	Billiard Ball Model
J.J. Thomson	1897	Negatively charged electron.	Plum Pudding Model
Niels Bohr	1913	Electrons move fixed orbits around the atom's nucleus in specific layers, or electron shells . Atoms absorb or give off energy when the electrons move from one shell (energy level) to another.	Bohr Model
Ernest Rutherford	1909	1909 (Based on idea that positive reflects positive) atom mostly empty space; electrons orbit randomly around positively charged nucleus.	Planetary Model

4. Periodic Table

- a. The current periodic table is arranged by increasing **atomic number**.
- b. Reading the Periodic Table
- An element's **properties** can be predicted from its location in the periodic table.
 - The atomic number represents the number of **protons**.
 - The atomic mass represents the number of **protons and neutrons**.

2nd Semester Final Exam Study Guide Answers

**Linda Lane Lilley, Shelly Rainforth
Collins, Julie S. Snyder**

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Study Guide for Pharmacology and the Nursing Process Linda Lane Lilley, Julie S. Snyder, Shelly Rainforth Collins, 2012-12-17

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one alternate format question per chapter An increased focus on prioritization provides practice in identifying the most important need to know nursing actions Drug dosage calculation questions one of which appears in every chapter facilitate mastery of the mathematics of drug dosing [Directory of Distance Learning Opportunities](#) Modoc Press, Inc.,2003-02-28 This book provides an overview of current K 12 courses and programs offered in the United States as correspondence study or via such electronic delivery systems as satellite cable or the Internet The Directory includes over 6 000 courses offered by 154 institutions or distance learning consortium members Following an introduction that describes existing practices and delivery methods the Directory offers three indexes Subject Index of Courses Offered by Level Course Level Index Geographic Index All information was supplied by the institutions Entries include current contact information a description of the institution and the courses offered grade level and admission information tuition and fee information enrollment periods delivery information equipment requirements credit and grading information library services and accreditation

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Gould,2024-05-31 After years of preparation and anticipation many students arrive at college without any real knowledge of the ins and outs of college life They ve been focused on finding the right school and have been carefully guided through the nuances of the admissions process but too often they have little knowledge about how college will be different from high school or what will be expected of them during that crucial first year and beyond Written by an award winning teacher How to Succeed in College While Really Trying provides much needed help to students offering practical tips and specific study strategies that will equip them to excel in their new environment Drawing on years of experience teaching at a variety of

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