

Chemistry 2nd Semester Final Review **KEY**

This is meant as a guide. It does NOT contain every single thing you need to know for the final exam. Review notes, benchmark review sheets, homeworks, chapter study guides, etc.

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Gas Laws (Ch. 13):

Subtopics: atmospheric pressure, Dalton's law of partial pressures, combined gas law, Ideal gas law, gas stoichiometry (22.4 L/mol for a gas at STP)

Practice questions:

1. If a gas occupies 3.8 L at a pressure of 2.71 atm, what would the pressure be if the volume changes to 1.47 L? (Temp. and amt. of gas constant)
V1=3.8 L; P1=2.71 atm
V2=1.47L; P2=? Atm
Since all units are standard (L, atm), we do not need to convert the units before proceeding.
 $(P_1V_1)/T_1=(P_2V_2)/T_2$
T goes away (not mentioned in problem)
 $P_1V_1=P_2V_2$
Isolate for P2: $P_2=(P_1V_1)/V_2$
 $P_2=(2.71 \text{ atm} \times 3.8 \text{ L})/(1.47 \text{ L})= 7.0 \text{ atm}$ (L cancelled and it's 2 sig figs)
2. A gas has a volume of 1.49 L at a temperature of 34.75 degrees Celsius. What would the volume be at 78.41 degrees Celsius? (pressure & amt. of gas constant)
V1= 1.49 L, T1= 34.75 degrees Celsius
V2= ? L, T2= 78.41 degrees Celsius
We need to convert from Celsius to Kelvins before proceeding.
T1= 34.75 + 273= 307.75 K
T2= 78.41 + 273= 351.41 K
 $(P_1V_1)/T_1=(P_2V_2)/T_2$
Get rid of P
 $V_1/T_1= V_2/T_2$
Isolate V2: $V_2=(V_1/T_1)T_2$
 $V_2=(1.49 \text{ L}/307.75 \text{ K})(351.41 \text{ K})= 1.70 \text{ L}$
3. What volume is occupied by 8.47 g of hydrogen gas at 84.7 degrees Celsius and 1.04 atm?
The combined gas law won't work here. You only have one condition and nothing changes.
PV=nRT will work. (n= moles and R= 0.08206 L atm/K mol)
P= 1.04 atm, V= ? L, n= OOPS we need to convert from grams to moles!, T= 84.7 degrees Celsius + 273= 357.7 K
n: Hydrogen gas is diatomic!
 $8.47 \text{ g H}_2 \times (1 \text{ mol H}_2/ 2.02 \text{ g H}_2)= 4.193069307 \text{ mol H}_2$
Isolate V: $V=(nRT)/P=(4.193069307 \text{ mol H}_2 \times 0.08206 \text{ L atm/ K mol} \times 357.7 \text{ K})/ 1.04 \text{ atm}= 118 \text{ L}$
4. What volume is occupied by 56.75 g of oxygen gas at STP?

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Minjie Lin



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