

**Physical Science – 2nd Semester – Final Exam Study Guide – (2012–2013)****1. Physical Science Basics**

- a. What tool(s) would you want to use to find the:
- |  |  |
|--|--|
| i. Mass of an object? <b>triple beam balance</b>                   | Basic SI Unit? <b>kg</b>                   |
| ii. Volume of an object? <b>ruler or water displacement method</b> | Basic SI Unit? <b>cm<sup>3</sup> or mL</b> |
| iii. Length of an object? <b>Ruler</b>                             | Basic SI Unit? <b>meter</b>                |
| iv. Density of an object? <b>graduated cylinder</b>                | Basic SI Unit? <b>mL</b>                   |
- b. What is the formula for calculating density? **mass/volume**
- i. What is the density of a cube that is 35 g and has a volume of 3 cm<sup>3</sup>? **35/3cm<sup>3</sup>**
- ii. Compare the density of solid H<sub>2</sub>O to liquid H<sub>2</sub>O, and explain what observable effect this has.  
**solid H<sub>2</sub>O less dense than liquid H<sub>2</sub>O**

**2. Chemistry Basics**

- a. Element (def.) **substance made up of one kind of atom**
- b. Atom (def.) **basic building block of matter**
- c. Compound (def.) **a molecule that contains atoms or molecules of two or more different elements**
- d. Molecule (def.) **a combination of two or more of the same or different atoms that are bonded together**

**3. Model of the Atomic Evolution of the Atomic Theory**

Atoms are so small that many scientists throughout history have created **models** to describe them.

Why have atomic theories changed over time? **better technology came around.**

Scientist	Year	Major Discovery/ Contribution	Name of atomic model
John Dalton	1808	Each element is made up of small, indivisible particles called atoms. Different elements have atoms of different masses.	Billiard Ball Model
J.J. Thomson	1897	Negatively charged electron.	Plum Pudding Model
Niels Bohr	1913	Electrons move <b>fixed orbits</b> around the atom's nucleus in specific layers, or <b>electron shells</b> . Atoms absorb or give off energy when the electrons move from one shell ( <b>energy level</b> ) to another.	Bohr Model
Ernest Rutherford	1909	1909 (Based on idea that positive reflects positive) atom mostly empty space; electrons orbit randomly around positively charged nucleus.	Planetary Model

**4. Periodic Table**

- a. The current periodic table is arranged by increasing **atomic number**.
- b. Reading the Periodic Table
- i. An element's **properties** can be predicted from its location in the periodic table.
- ii. The atomic number represents the number of **protons**.
- iii. The atomic mass represents the number of **protons and neutrons**.

# 2nd Semester Final Exam Study Guide Answers

**Peter L. Kresan, Reed Mencke**



**2nd Semester Final Exam Study Guide Answers:**

## The Enigmatic Realm of **2nd Semester Final Exam Study Guide Answers**: Unleashing the Language is Inner Magic

In a fast-paced digital era where connections and knowledge intertwine, the enigmatic realm of language reveals its inherent magic. Its capacity to stir emotions, ignite contemplation, and catalyze profound transformations is nothing short of extraordinary. Within the captivating pages of **2nd Semester Final Exam Study Guide Answers** a literary masterpiece penned by a renowned author, readers attempt a transformative journey, unlocking the secrets and untapped potential embedded within each word. In this evaluation, we shall explore the book's core themes, assess its distinct writing style, and delve into its lasting effect on the hearts and minds of those who partake in its reading experience.

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