

18.4 Molecular Structure: Mass

Multiple Choice

1. The molecular formula of a compound is the chemical formula written _____
 (A) percent composition (B) percent yield
 (C) molar formula (D) structural formula
2. Percent of hydrogen in $\text{C}_2\text{H}_6\text{O}$ is _____
 (A) 15.0% (B) 12.5%

3. In a gas the other reactions represent oxygen mixed with nitrogen, 2.5 g of gas was produced by the reaction, and only produced 1.5 g of gas. What is the percent yield of sulfur in this reaction?



(A) 60% (B) 75% (C) 80% (D) 90% (E) 100%

4. In this reaction $\text{ZnCO}_3 \rightarrow \text{ZnO} + \text{CO}_2$, what is the percent yield of oxygen?
 (A) 100% (B) 80% (C) 70% (D) 50%

5. What is the percent yield of oxygen in the following reaction?



(A) 100% (B) 75% (C) 50% (D) 25%

6. Aluminum reacts with H_2 gas to produce hydrogen gas. If 10.0 g of aluminum reacts with H_2 gas, what mass of H_2 gas is produced?



(A) 10.0 g (B) 15.0 g (C) 20.0 g (D) 30.0 g (E) 45.0 g

7. In a full combustion, structural formula of $\text{C}_2\text{H}_6\text{O}$ is what gas is produced? $\text{C}_2\text{H}_6\text{O} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$. What is the percent yield of CO_2 in this reaction?



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