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web 1 molarity mv mass molar mass x 0 4500 l 0 825 g 141 9579 g mol x 0 0129 m 2 molality 0 825 g 141 9579 g mol 0 00581158 mol 0 00581158 mol 0 4500 kg 0 0129 m 3 mole fraction na 2 hpo 4 0 825 g 141 9579 g mol 0 00581158 mol h 2 o 450 0 g 18 015 g mol 24 97918401 mol

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web apr 7 2014 molarity moles of solute liters of solution 8 4 2 2 a first convert 250 ml to liters 250 1000 0 25 then

calculate molarity 5 moles 0.25 liters 20 m³ c a solution with molarity 2 requires 2 m of n a oh per liter so 4 x 2.8 m⁴ a a solution of molarity 1.5 m requires 1.5 mol of na to every litre of solvent

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$$m_{mix} = \frac{m_1 v_1 + m_2 v_2 + \dots + m_n v_n}{v_1 + v_2 + \dots + v_n}$$
 1.5 x 480 + 1.2 x 520 / 480 + 520 = 720 / 1000 = 0.72 m
 with increase in temperature which of these changes a molality b fraction of solute present in water

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